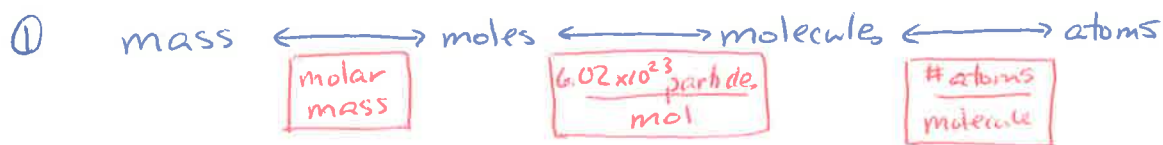


## UNIT 2 OVERVIEW



② % comp. A by mass =  $\frac{\text{molar mass element A}}{\text{molar mass compound}} \times 100\%$

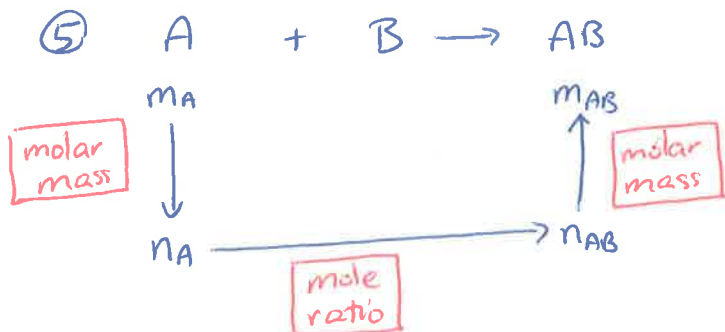
③a Empirical formula from % comp.



$\uparrow$  may be given these directly eg. hydrates

3b ④ Molecular Formula

$$\frac{\text{molar mass}}{\text{empirical formula mass}} = \text{multiplier}$$



⑥ Limiting/Excess Reagent - calculation #'s to other reactant.  
- compare calculated value to given value.

⑦ % yield =  $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100\%$